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Acetate Trihydrate Crystals Melt At  $136.4\text{ }^\circ\text{F}/58\text{ }^\circ\text{C}$  (to  $137.12\text{ }^\circ\text{F}/58.4\text{ }^\circ\text{C}$ ), Dissolving  
In Their Water Of Crystallization. When They Are Heated Past The Melting Point And  
Subsequently Allowed To Cool, The Aqueous Solution Becomes Supersaturated. This  
Solution Apr 15th, 2024

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67 PREPARATION AND TESTING OF BUFFER SOLUTIONS. P. URPOSE. The Purpose Of The Laboratory Investigation Is To Experimentally Determine (1) PKa (and Thus Ka) Of The Acid In A Buffer And Thus The Buffer Range, (2) Investigate The Buffer Capacity Of Feb 3th, 2024

### **Section 19.1. Acid-Base Buffer Solutions**

If We Add CH<sub>3</sub>COO-SHIFT Net Effect: [H<sup>30+</sup>] Decreases PH Increases Also, [OH<sup>-</sup>]

Increases POH Decreases I.e., a Weak Acid And Its Conjugate Base Or A Weak Base And Its Conjugate Acid. 19-2 Example: Addition Of  $\text{NH}_4\text{Cl}$  To  $\text{NH}_3$  Jan 24th, 2024

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