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That Contains 6.2×10^{-3} Mole Of HCl In 2.2 L Of Solution.
 $\text{HCl} \rightarrow \text{H}^+ + \text{Cl}^-$ Molarity Of HCl (which Is The H^+ Same As H_3O^+)
 $= \frac{6.2 \times 10^{-3} \text{ Mol}}{2.2 \text{ L}} = 2.9 \times 10^{-3} \text{ M}$ [H 3 Feb 6th, 2024
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