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Answer On Question #60099, Chemistry / Inorganic Chemistry 1. 1.00g Of Carbon Is Combusted In A Limited Supply Of Pure Oxygen. 0.50g Of The Carbon Combusts To Form CO<sub>2</sub> And 0.50 G Of The Carbon Combusts To Form CO. The Resultant Mixture Of CO<sub>2</sub> And CO Is Passed Through Excess Na May 1th, 2024.

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Produces 0.0322 Mol Of Acetanilide. The Amount Of Acetic Anhydride Does Not Play A Role, Since It Is In Excess. Then,  $n(\text{N}) = 0.0322$  .aniline  $n(\text{Aceta II}) = \text{Ide Mol}$  We Find The Molar Mass Of Acetanilide:  $(\text{N}) (\text{O}) = 128.16 + 135(1.89) = 250.82$  M Aceta ... Mar 3th, 2024

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