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 KEY Chemistry 11 Stoichiometry Review Worksheet Name: KEY 33. $\text{KClO}_4 \rightarrow 2\text{O}_2 + \text{KCl}$ Actual yield = 1.45g Theoretical yield = 5.95g $\text{KClO}_4 \times 1 \text{ mol} \times 122.55 \text{ g} \times 20 \times 2$
 $1 \text{ KClO}_4 \times 32.00 \text{ g} \times 1 \text{ mol} = 3.11 \text{ g} \times 2 \text{ } \% \text{ Yield} = \frac{1.45 \text{ g}}{3.11 \text{ g}} \times 100\% = 46.6\%$ 34. $4\text{FeCO}_3 + \text{O}_2 \rightarrow 2\text{Fe}_2\text{O}_3 + 4\text{CO}_2$ Percent purity = 62.8%? g pure FeCO_3 Jan 5th, 2024.
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 Of LiNO_3) 4) Using The Following Equation: $\text{Fe}_2\text{O}_3 + 3\text{H}_2 \rightarrow 2\text{Fe} + 3\text{H}_2\text{O}$.
 Calculate How Many Grams Of Iron Can Be Made From 16.5 Grams Of Fe_2O_3 By The
 Following Equation. Feb 13th, 2024 Basic Stoichiometry Phet Post Lab Answer
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Answer Key - Bing Phet Simulation The Skater Had More Potential Energy At The Bottom And Had More Kinetic Energy At The Top. The Relationship Between Kinetic And Potential Energy Is That When One Goes Up, The Other Comes Down.

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