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$\Delta G = \Delta G^0 + RT \ln \frac{a_{H^+} a_{O_2}}{a_{H_2}}$ (2.1.36) The Free Energy Of The Cell Reaction Is Given By, $\Delta G = \Delta G^0 + RT \ln \frac{a_{R} a_{A} a_{N} a_{H^+} a_{O_2}}{a_{H_2}}$ (2.1.38) Which Since $\Delta G = -nFE_{rxn}$ And $a_{H^+} = a_{H_2} = 1$ Becomes, $E = E^0 + \frac{RT}{nF} \ln \frac{a_{R} a_{A} a_{O_2}}{a_{N}}$. (2.1.40) This Expression Is Called The Nernst Equation. This Relation Can Jan 14th, 2024

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