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 $E = E^0 - 0.0592 \text{ N} \log Q$

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$$2 \text{ H}_2(\text{g}) + \text{O}_2(\text{g}) \rightarrow 2 \text{ H}_2\text{O}(\text{l})$$

In A Hydrogen-oxygen Fuel Cell, Energy Is Produced By The Overall Reaction Represented Above. (a) When The Fuel Cell Operates At 25°C And 1.00 Atm For 78.0 Minutes, 0.0746 Mol Of O

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