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 $K\cdot\text{mol}$. If Pressure Is Needed In KPa Then Convert By Multiplying By $101.3 \text{ kPa} / 1 \text{ atm}$ to Get. $R = 8.31 \text{ L}\cdot\text{kPa} / (\text{K}\cdot\text{mole})$ | May 7th, 2024
Chapters 10 & 11 - Gases, Gas Laws, And Gas Stoichiometry ...
Graham's Law Ideal Gas Law Molar Volume Molecular Mass Determination (of A Gas)
OBJECTIVES: • Memorize The Values For STP. • Memorize And Be Able To Apply The Gas Laws: Boyle's, Charles, Dalton's Law Of Partial Pressure, Combined Gas Law, Gay-Lussac's, And Graham's. • Be Able To Use M
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Molecular Picture Of Gas Gas Is Made Up Of Many Individual Molecules
Number Density Is Number Of Molecules/volume: $N/V = R/m$ R Is The Mass Density M Is The Mass For One Molecule
Number Of Moles: $N = N / N_A$ N_A Feb 21th, 2024

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