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Chlorine F. Neese Et Al., Inorg. Chem., 1999, 38, 4854. Transition Dipole Integral  $3p$   
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The Equilibrium Constant In Either Of Two Ways:  $K_{\text{association}} = K_a = \frac{[R] \cdot [L]}{[RL]}$   
 $K_{\text{dissociation}} = K_d = \frac{[RL]}{[R] \cdot [L]}$  Where The Concentrations Of The Free R, Free L  
And The Complex Are The Concentrations At Equilibrium  $K_a = 1/K_d$  It Is Also Easy To  
Show That  $G \rightleftharpoons O$  Association = -  $G \rightleftharpoons O$  Dissociation May 2th, 2024 Generative

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