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Name: KEY 33. $\text{KClO}_4 \rightarrow 2\text{O}_2 + \text{KCl}$

Actual yield = 1.45g Theoretical yield = 5.95g $\text{KClO}_4 \times 1\text{mol}$

$122.55\text{g} \times 2\text{O}_2 \times 32.00\text{g} \times 1\text{mol} = 3.11\text{g O}_2$

%Yield = $1.45\text{g} \times 3.11\text{g} \times 100\% = 46.6\%$ 34. $4\text{FeCO}_3 + \text{O}_2$

$\rightarrow 2\text{Fe}_2\text{O}_3 + 4\text{CO}_2$

Percent purity = 62.8%? pure FeCO_3 Jan 3th, 2024

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Molecules Should Be Supplied To Burn 1 Mol Of CH_4 ...

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Answers Calcium $\text{C}_2\text{O}_2 + 3\text{Cr}(3)$ When The Potassium

Chlorate (KClO_3) Is Heated, It Decomposes To Form

Gases KCl And Oxygen (O_2 .) $2\text{KClO}_3 \rightarrow 2\text{KCl} + 3\text{O}_2$ C.

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