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Theoretical yield = $5.95\text{g KClO}_4 \times \frac{1\text{mol}}{122.55\text{g}} \times 2 \times 32.00\text{g/mol} = 3.11\text{g O}_2$ %Yield = $\frac{1.45\text{g}}{3.11\text{g}} \times 100\% = 46.6\%$ 34.
 $4\text{FeCO}_3 + \text{O}_2 \rightarrow 2\text{Fe}_2\text{O}_3 + 4\text{CO}_2$ Percent purity = 62.8%? g pure FeCO_3 Apr 14th, 2024

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