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The yield = $5.95 \text{ g KClO}_4 \times \frac{1 \text{ mol}}{122.55 \text{ g}} \times 2 \text{ O}_2 \times \frac{1 \text{ KClO}_4}{4 \text{ O}_2} \times 32.00 \text{ g/mol} = 3.11 \text{ g O}_2$
% Yield = $\frac{1.45 \text{ g}}{3.11 \text{ g}} \times 100\% = 46.6\%$
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