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1. A Solution Is Prepared By Mixing 100.0 G Of Water,  
H<sub>2</sub>O, And 100.0 G Of Ethanol, C<sub>2</sub>H<sub>5</sub>OH. Determine The  
Mole Fractions Of Each Substance. 2. The Molality Of  
An Aqueous Solution Of Sugar (C<sub>12</sub>H<sub>22</sub>O<sub>11</sub>) Is 1.62m.  
Calculate The Mole Fractions Of Sugar And Water. 3.  
Calculate The Molality Of 25.0 Grams Of 3th, 2024

## **Mole Fraction Molality Molarity - University Of Texas At ...**

Mole Fraction Sugar Of Solution  $\chi_{\text{Sugar}} = (0.292$   
 $\text{Mol}) / (0.292 + 55.6) = .00522$  Mole Fraction Water Of  
Solution  $\chi_{\text{Water}} = (55.6 \text{ Mol}) / (0.292 + 55.6) = 0.995$

(or 1-  $\chi$  Sugar) Molality  $M = (.292 \text{ Mol})/(1 \text{ Kg}) = 0.292 \text{ Mol Kg}^{-1}$   
Molarity  $M = (.292 \text{ Mol})/(1 \text{ 3th, 2024})$

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3.mole Fraction = 0.950 4.mole Fraction = 0.0360  
5.mole Fraction = 0.925 Nsci140802exam 018  
10.0points 3.3 G Of NaCl (58.4 G/mol) And 9 G Of KBr (119.8 G/mol) Were Dissolved In 40 G Of Water (18.0 G/mol). What Is The Mole Fraction Of KBr In The Solution? Nsci140805 019 10.0points How Many G 2th, 2024

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As  $1 \text{ atm} = 1.013 \times 10^5 \text{ Pa}$ , The Osmotic Pressure Is  $\Pi = (0.0036 \times 1.013 \times 10^5) \text{ Pa} = 365 \text{ Pa}$  Hence  $C = \Pi / RT = (365 \text{ Pa}) / ((8.314 \text{ J K}^{-1} \text{ Mol}^{-1}) \times (293 \text{ K})) = 0.15 \text{ Mol M}^{-3}$  As  $1 \text{ M}^3 = 1000 \text{ L}$ , This Corresponds To  $C = 1.5 \times 10^{-4} \text{ M}$  1th, 2024

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