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Calculations From Chemical Equations Mole - Mole ... $7 + 6 \text{KI} + 7 \text{H}_2\text{SO}_4 \rightarrow \text{Cr}_2(\text{SO}_4)_3 + 4 \text{K}_2\text{SO}_4 + 3 \text{I}_2 + 7 \text{H}_2\text{O}$  A) How Many Moles Of Potassium Dichromate ( $\text{K}_2\text{Cr}_2\text{O}_7$ ) Are Required ... = 407.9 G AgBr This Is The Theoretical Yield Yields 22

B) Calculate The Percent Yield If 375.0 G Of Silver Bromide Was Obtained From The Reaction Theoretical Yield = 407.9 G AgBr Percent Yield =  $100 \times \frac{\text{Actual Yield}}{\text{Theoretical Yield}}$  May 6th, 2024 Stoichiometry Worksheet #2 (mole-mass, Mass-mole Problems)

Stoichiometry Worksheet #2 (mole-mass, Mass-mole Problems) 1.  $\text{N}_2 + 2\text{O}_2 \rightarrow \text{N}_2\text{O}_4$  A. If 15.0g Of  $\text{N}_2\text{O}_4$  Was Produced, How Many Moles Of  $\text{O}_2$  Were Required?  $15.0\text{g } \text{N}_2\text{O}_4 \times \frac{1 \text{ Mol } \text{N}_2\text{O}_4}{2 \text{ Mol } \text{O}_2} \times \frac{92.0\text{g } \text{N}_2\text{O}_4}{1 \text{ Mol } \text{N}_2\text{O}_4} = 0.326 \text{ Mol } \text{O}_2$  B. If  $4.0 \times 10^{-3}$  Moles Of Oxygen Reacted, How Many Grams Of  $\text{N}_2$  Were Needed?  $4.0 \times 10^{-3} \text{ Mol } \text{O}_2 \times \frac{1 \text{ Mol } \text{N}_2}{2 \text{ Mol } \text{O}_2} \times 28 \dots$  Mar 1th, 2024 CHEMISTRY

WORKSHEET # 2 MOLE PROBLEMS—THE MOLE ... CHEMISTRY WORKSHEET # 2: THE MOLE AS A UNIT OF MASS Define The Term Molar Mass (worksheet #1): \_\_\_\_\_ Now That You Know How To Find The Mass Of One Mole Of A Substance (molar Mass) You Can Easily Find The Mass Of Several Moles Or The Mass Of A Fraction Of A Mole Using The Factor-label Technique. Jan 7th, 2024.

Worksheet: Mixed Problems—Mole/Mole Name And ...  $2 \text{ Cu} + \text{O}_2 \rightarrow 2 \text{ CuO}$  A. If 101 Grams Of Copper Is Used, How Many Moles Of Copper (II) Oxide Will Be Formed? B. If 5.25 Moles Of Copper Are Used, How Many Moles Of Oxygen Must Also Be Used? C. If 78.2 Grams Of Oxygen React With Copper, How Many Moles Of Copper (II) Oxide Will Be Produced? 2.  $\text{C}_4\text{H}_{10} + \text{O}_2 \rightarrow \text{CO}_2 + \text{H}_2\text{O}$  A. How Many Moles Of

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