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HII/N H = 0.025 At T=12000 K, N HII/N H = 50.5 The Fraction Of Ionized Atoms Is N HII N Total = N HII (N H+ N HII) = N HII=N H 1 + N HII=N H (1) So At T=8000, Approximately 2% Of Hydrogen Atoms Are Ionized. Conversely, At T=12000K 98% Of Hydrogen Atoms Are Ionized. D) Clearly, Ionization Occurs At A Much Lower Temperature Than ... 2th, 2024Saha's Equation: Dissociation And Ionization Of Hydrogen D ...B. Dissociation And Ionization Of Hydrogen Now, The Above Formula Can Be Used As A Template For The Reactions Of Interest For This Write-up. The Dissociation Of Molecular Hydrogen And The Ionization Of Atomic Hydrogen: ! H 2 "2H"2p ++2e# Letting The Initial Concentration Of H 2 Be N 0, We Can Write Down The Concentration Of All 4th, 2024.

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