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And Thermodynamics . Course Schedule . N.B.: The Schedule Below Has Not Yet Been Adapted To The Blended Schedule Of Online And In-class Meetings. Course Readings May Vary Between Course Offerings.

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Basic Concepts Of Thermodynamics Absolute Pressure Gauge Pressure Vacuum Pressure Absolute Vacuum Pressure ABSOLUTE ATMOSPHERIC PRESSURE Pressure  $P$  Atm

Temperature • Temperature Is A Pointer For The Direction Of Energy Transfer As Heat  $Q$   $Q$   $T$   $A$   $T$   $A$   $T$   $A$   $T$   $A$   $T$   $B$   $T$   $B$   $T$   $B$   $T$   $B$  >

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Thermodynamics, The First Law: The Concepts  
The Internal Energy Is An Extensive Property - It Depends On The Amount Of Substance. The Molar Internal Energy,  $U_m = U/n$  - Intensive Property, Does Not Depend On The Amount Of Substance, But Depends On The Temperature And Pressure. Internal Energy, Heat, And Work Are All Measured 6th, 2024  
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The Molar Specific Volume Is Then An Intensive Property. 2-8C A Process During Which A System Remains Almost In Equilibrium At All Times Is Called A Quasi-equilibrium Process. Many Engineering 8th, 2024.

3. Basic Concepts Of Thermodynamics - Part 2  
At Constant Volume The Heat Absorbed Increases The Energy Of The System When The Temperature Is Raised From  $T_1$  To  $T_2$ , I.e.  $C_V(T_2 - T_1) = U_2 - U_1$ .  
2 For A Very Small Change  $dT$  In Temperature, The Heat Capacity At Constant Volume Is Equal To The Rate Of Change  
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